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1 - AP Chemistry

Chemical Equations

Worksheet Write the

balanced chemical

equation (excluding

spectator ions)

underneath each

reaction description,

and answer the

question.

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## **10826 AP Chemistry Chemical Equations Worksheet**

Answer: Hydrogen gas is released at the cathode (the negative electrode). (o) A

solution of copper(II) sulfate is electrolyzed using inert electrodes.

$2\text{Cu}^{2+} + 2\text{H}_2\text{O} \rightarrow 2\text{Cu} + \text{O}_2 + 4\text{H}^+$  Is the final solution acidic, basic, or neutral?

Answer: Acidic hydrogen ions ( $\text{H}^+$ ) are produced in the

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reaction. (p) A solution of ammonium sulfate is added to a satu-

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continued 2 21 linn  
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Chemistry of Gases  
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Supplementary  
Chemical Equations

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Worksheet Predict the products and balance the following chemical equations to identify the contestants in the

Chemistry of Gases tournament. 1.Na 2  
CO 3 (s) + \_\_\_HCl(aq)→

## **Sweet 16 Chemistry of Gases**

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questions in Chapter 8:  
ROUND 3 Exercise 8-1:  
Using the activity  
series, predict and  
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single replacement  
reactions.

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## Equations Flinn **Chemistry Chemical Equations Flinn Answer Key**

This is the most common type of chemical equation. In this chemical equation, a new product is formed by combining two to three combinations of reactants. For instance,  $H_2 + O_2 \rightarrow H_2O$ . This is a chemical equation where two atoms of

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hydrogen are combined to form a product, water. This is why this reaction is called as synthesis reaction.

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Replace immutable groups in compounds to avoid ambiguity. For example,  $C_6H_5C_2H_5 + O_2 = C_6H_5OH + CO_2 + H_2O$  will not be

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balanced, but  $\text{XC}_2\text{H}_5 + \text{O}_2 = \text{XOH} + \text{CO}_2 + \text{H}_2\text{O}$  will. Compound states [like (s) (aq) or (g)] are not required. You can use parenthesis () or brackets [].

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## **Flinn Scientific**

1 CH<sub>4</sub> + 2 O<sub>2</sub> → 1 CO<sub>2</sub>

+ 2 H<sub>2</sub>O 2. 2 Na<sup>+</sup>

+ 2 Cl<sup>-</sup> → 1 NaCl 3. 4 Al + 3

O<sub>2</sub>

**Name: Date:**

## **Balancing Equations**

Every balanced

chemical equation

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consists of two parts: the reactant side and the product side. Both of these sides are separated by the means of an arrow. On the left side of the arrow, you will find the reactant side. This side represents the elements which are used for initiating the experiment.

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Substitute immutable groups in chemical compounds to avoid ambiguity. For instance equation  $C_6H_5C_2H_5 + O_2 = C_6H_5OH + CO_2 + H_2O$  will not be balanced, but  $PhC_2H_5 + O_2 = PhOH + CO_2 + H_2O$  will; Compound states [like (s) (aq) or (g)] are not required. If you do not know what products are enter reagents only and click 'Balance'. In many

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cases a complete equation will be suggested.

## **Balance Chemical Equation - Online Balancer**

Chemical reactions have the same number of atoms before the reaction as after the reaction. Balancing chemical equations is a basic skill in chemistry and testing yourself helps retain important information. This

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collection of ten  
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questions will give you  
practice in how to  
balance chemical  
reactions.

## **Balancing Equations Chemistry Test Questions**

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and Reactions MIXED  
REVIEW SHORT

ANSWER Answer the  
following questions in  
the space provided. 1.  
b A balanced chemical

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equation represents all the following except (a) experimentally established facts. (b) the mechanism by which reactants combine to form products. (c) identities of reactants and products in a chemical reaction.

## **8 Chemical Equations and Reactions**

The chemical equation for this reaction is

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written as:  $2\text{H}_2 + \text{O}_2$

$\rightarrow 2\text{H}_2\text{O}$  Chemical

formulas and other symbols are used to indicate the starting materials, or reactants, which by convention are written on the left side of the equation, and the final compounds, or products, which are written on the right.

## **3.1: Chemical Equations - Chemistry**

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To get four hydrogen atoms on the right, add a coefficient of 2 for the hydrogen gas. The coefficient is a number that goes in front of a chemical formula.

Remember, coefficients are multipliers, so if we write  $2 \text{H}_2\text{O}$  it denotes  $2 \times 2 = 4$  hydrogen atoms and  $2 \times 1 = 2$  oxygen atoms.

$$\text{SnO}_2 + 2 \text{H}_2 \rightarrow \text{Sn} + 2 \text{H}_2\text{O}$$

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and hydrochloric acid.  
In this lab, students ...

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'Write a balanced  
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